

THE INFLUENCE OF EXTRANEEOUS ELEMENTS ON
LINE INTENSITIES IN THE A.C. SPARK AND THE D.C. ARC

by

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INTRODUCTION

During the past few years considerable attention has been directed toward the influence of extraneous elements on line intensities in spectrochemical work. This constitutes a major problem in trace mineral investigations, since these influences are more pronounced if the elements under consideration are in low concentrations. These influences have been investigated by numerous scientists (4,6,7,8,9,14,15,17).

In this research the influence of calcium upon the intensity of magnesium both in the a.c. spark and the d.c. arc was investigated. The procedure was then reversed to study the influence of magnesium upon calcium in the two sources. These two elements are present as minerals in plants, so it was thought advisable to use concentrations approximating those present in the common forage plants. It is well known that the effect of extraneous elements can be decreased or eliminated through the use of spectrographic buffers. Therefore, the same procedures were repeated except that a buffer was added to the samples being investigated.

The investigators listed above have studied many sequences of elements under varying conditions. Some of the elements which have been studied are: Na, V, Pb, Sn, Mn, Fe, Al, Bi, Ca, Cu, Zn, Ti, P, K, Li, Cd, Rb, Si, Cr, Ni, Zr, La, Ag, Mo, Ga, C, and B. However, there has been little work done on correlating the effects of different excitation sources on the influences of

extraneous elements. The results obtained by these workers have not always been in close agreement. The combination of elements investigated in this research has not been reported by other workers. If the effect of one element upon another can be determined and nullified, it would lead to higher accuracy in trace mineral analyses by spectrochemical means.

SURVEY OF LITERATURE

Brode and Silverthorn (8) through their study on the effects of impurities on line intensities of Cd and Zn in quantitative spectrochemical analysis have shown the following results. Copper and bismuth had an appreciable effect on the relative intensities of the cadmium and zinc lines. The effects were in opposite directions for in the presence of copper, the intensity of the zinc lines was reduced as compared to cadmium, and bismuth reduced the intensity of the cadmium line compared with zinc lines.

The intensities of spectral lines of vanadium have been increased through the addition of sodium chloride (4). This enhancement was noticeable for lines of both the neutral and ionized atom. However, potassium chloride intensified the lines of the neutral atom and decreased the intensity of the lines of the ionized atom.

Wada (16) investigated the reciprocal effects of the alkali metals on their detection in the electric arc. The following

conclusions were drawn from his study: (a) the presence of potassium increased the sensitivity of the sodium test; (b) the test for potassium was not affected by the presence of sodium and lithium, but it increased the sensitivity of tests for all other alkalies. The results of his work were illustrated by plotted curves.

Waldbauer and Means (17) have shown that the intensity of the spectral lines of calcium increased with an increase in the percentage of the alkali metal present. However, at higher concentrations the line enhancement tended to be decreased.

Brode and Hodge (7), and Brode and Bournique (6) investigated the influence of extraneous anions on line intensities. They concluded that in a given series of compounds of different anions but with common cations, the line intensity of the cation is greatest for those compounds having the lower boiling point.

Duffendack, Wiley, and Owens (10) concluded that the influence of the extraneous element on the ratios of the intensity of the element under investigation to the internal standard reached a maximum value. This maximum influence was attained through the addition of a large amount of the extraneous element to the standard solutions and unknowns.

It has been indicated, through the work of Brode and Timma (9), that the effect of the extraneous element varies with the amount of the extraneous element, but not in a simple linear fashion. They were able to arrange a sequence of elements

based on their boiling points and excitation potentials in which the intensities of the elements above a given one were enhanced and those below were decreased. This sequence agreed closely with a mathematical expression which included the boiling point and excitation potential of the element. However, if such an arrangement of elements were studied, it would require a strict reproducibility of the excitation sources used in their work.

Marked variations have been observed in the relative intensities of magnesium, cadmium, and lead lines when amounts of foreign substances were added to standard samples and excited in a condensed spark source (13). The changes in the interspectra intensity ratios depended on the nature and amount of the added substances, as well as, on the elements under consideration.

Langstroth and McRae (12) stated in their work that unless appropriate precautions were taken, the intensity ratios and therefore the determined concentrations, varied with variations in the extraneous composition of the samples, even though the amounts of the internal standard and the investigated element present remained constant. The authors indicated the necessity for the use of a spectroscopic buffer and provided certain criteria for the choice of the buffer and internal standard elements.

The effect of changing the conditions of the excitation source upon the influence of extraneous materials in spectrochemical analysis has been studied by Andrychuck and

McMullen (2). They reported that the extraneous material affected the mechanism of volatilization more than the mechanism of excitation. Operating the excitation source at a temperature above the boiling points of the constituents of the sample and at high excitation energies as obtained from a low inductance spark source reduced the influence of extraneous materials.

Schrenk and Clements (15) have shown that potassium and calcium decreased the line intensity of phosphorus. Also, when cobalt was used as an internal standard, increasing amounts of calcium tended to enhance the cobalt line. It was noted that phosphorus increased the intensity of the cobalt line. Through the use of a lithium tartrate buffer, they were able to minimize these extraneous element effects.

EXPERIMENTAL PROCEDURE

Instruments Used

A Bausch and Lomb, large, Littrow spectrograph was used throughout this investigation. The instrument is provided with interchangeable quartz and glass optical systems, and in this research the quartz system was used. The quartz prism has a height of 57 mm and a refracting face of 95 mm, which utilizes the full aperture of the lens. The lens is 70 mm in diameter and has a focal length of 1827 mm.

Two positions were used in this investigation, position

five which covers the range of 2600 to 3350 Å, and position three which covers the range of 2900 to 4800 Å. These different positions are obtained by rotation of the prism, movement of the prism and lens to a new focus, and a change in the tilt of the photographic plate. The first two adjustments are coordinated by means of a cam, which causes the prism to rotate as it is moved forward or backward during focusing. The plate tilt is adjusted directly through a screw and recording scale. The dispersion of the instrument is 3.15 Å per mm at 2780 Å and 5.15 Å per mm at 3160 Å. The light from the source was focused with a condensing lens upon a cylindrical lens which then focused it upon the slit. A rotating adjustable sector on a motor mount was secured just ahead of the cylindrical lens. This sector was used to adjust exposure conditions for the different sources. It was also possible, through the use of this sector, to expose for sufficient time to make certain that all the sample was volatilized.

Two energy sources were used in conducting this research. The first source employed was a Bausch and Lomb condensed spark unit which had a fixed capacitance of 0.005 micro-farads and six steps for selection of inductance in the circuit. The inductance could be varied from zero to 20.56 micro-henrys. A transformer with a 0.45 KVA output furnished a secondary voltage of 15,000 volts a.c. from a primary voltage of 115 volts. The second source of excitation was a direct current generator maintained at a potential of 125 volts and a current of eight to ten amperes.

The following conditions were maintained constant: slit width, 60 microns; slit height, 1.0 mm; electrode separation, 2.5 mm; and exposure time, 60 seconds. When the a.c. spark excitation source was used the rotating sector was set at three-fourths, and when the direct current source was used the sector was set at one-eighth.

Kodak Number 33 photographic plates were found to give good results for the magnesium and calcium under investigation. All plates were developed four minutes at 20° C. in Kodak X-ray developer diluted with two parts of water. The plates were then placed in a fixing bath for ten minutes after which they were washed for thirty minutes in running water, and dried for eight minutes on a warm air drier. The transmittancies of the lines were measured with a projection type Allied Research Laboratories - Dietert densitometer. It was possible to obtain check readings of the transmittancies within a fraction of one percent.

Choice of the Calcium and Magnesium Lines

Calcium exhibits numerous spectral lines which may be adaptable to spectrochemical use. The line chosen for this research was the one having a wavelength of 3158.9 ⁰Å. According to Brode's Tables (5) this is a first ionization state line, having a spark intensity of 300 and an arc intensity of 100. The line is also relatively free of interference lines from other elements. Ahrens (1) states that the detection limit for calcium is

approximately 0.00005 percent. In using solutions compounded from analytical grade reagents, interfering elements were not a factor in the selection of the spectral line.

The element magnesium is extremely sensitive for spectrographic analyses. Using the most sensitive line of magnesium (2852.1 \AA), concentrations as low as 0.0001 percent are usually detectable (1). The line utilized in the magnesium investigation was the one having a wave length of 2779.8 \AA . Brode (5) lists this line as having a spark intensity of 50, an arc intensity of 40, and arising from a first ionization state. There are no interfering lines listed in Brode's Tables for this wave length.

Preparation of Samples

Analytical grade calcium carbonate was dissolved in a small portion of dilute hydrochloric acid and made up to volume with distilled water, so that the concentration of calcium was 2000 ppm. For magnesium analyses, stock solutions containing 2000 ppm of magnesium were made from analytical grade magnesium nitrate hexahydrate dissolved in water and made to volume. Table 1 shows the composition of the solutions used for the analyses.

For the buffered solutions, two ml of a lithium tartrate solution containing 81,000 ppm of lithium tartrate were added to each of the solutions mentioned in Table 1. The concentration of the buffer in the sample was then approximately 3860 ppm.

Table 1. Formulation of samples for analysis from stock solutions.

Solution : number	ml stock : Ca solution	ml stock : Mg solution	ml : H ₂ O	ppm : Ca	ppm : Mg
1	5	0	35	250	0
2	5	2	33	250	100
3	5	4	31	250	200
4	5	10	25	250	500
5	5	20	15	250	1000
1	0	5	35	0	250
2	2	5	33	100	250
3	4	5	31	200	250
4	10	5	25	500	250
5	20	5	15	1000	250

Preparation of Electrodes

Porous cup type electrodes are required when the a.c. spark excitation source is used. This type electrode was prepared from the standard 5/16 inch diameter carbons supplied by the National Carbon Company. The 5/16 inch diameter rods were cut into 1¼ inch lengths. These cut rods were then placed in a jig similar to the one described by Feldman (11) and drilled with a 5/32 inch drill through the center, along the axis, to within approximately 0.5 mm of the bottom. This gave a uniform thickness to the porous cup electrodes and the sample feed rate appeared to be constant throughout. The porous cup electrode served as the upper electrode and a 3/16 inch diameter carbon rod served as the counter electrode. Figure 1 shows a cross section of the porous cup electrode and also the counter electrode space indicated.

In utilizing the d.c. arc source, cupped type electrodes were used. The 5/16 inch diameter rods were cut into 1¼ inch lengths. The ends of the anode were crater shaped 1.5 mm deep. The cathodes, of the same diameter, were bullet shaped. On the end of the blunt nose of the cathode there was a cylindrically shaped projection 1 mm in diameter and 6 mm long. This projection tended to prevent wandering of the arc and also allowed a 60 second exposure without the electrode being consumed. The anode craters were filled with a saturated solution of carnauba



Fig. 1. The Porous Cup Electrode; with
the Counter Electrode

wax in carbon tetrachloride and dried for 45 minutes at 100° C. in an oven. Figure 2 shows the electrodes used in the d.c. arc.

Introduction of Samples into the Porous Cup Electrodes and onto the Cratered Electrodes

The porous cup electrodes were pre-sparked for 120 seconds when being used with an unbuffered solution. When calcium was being investigated, a 60 second pre-spark period was recorded on the photographic plate, and the intensity of the calcium line from the electrode was utilized as a correction factor for the calcium line of the sample. Pre-sparking was required for the solution to pass through the electrode uniformly. The solution samples were placed in the porous cup electrodes by means of a long, slender nosed dropper inserted to the bottom. The dropper was slowly removed as the electrode filled. This procedure of filling was necessary in order to prevent air from being trapped in the bottom of the electrode. After each sample was sparked, the dropper was rinsed at least five times with distilled water and twice with the next sample before loading another electrode. When working with a buffered solution in the a.c. spark, the solution would not pass through the electrode if it had been pre-sparked as indicated. However, the solution would pass through if the electrode had been pre-arc'd for twenty-five seconds. When calcium was under consideration, two pre-arc'd electrodes were sparked for sixty seconds on each plate to determine the calcium correction factor for the electrodes.

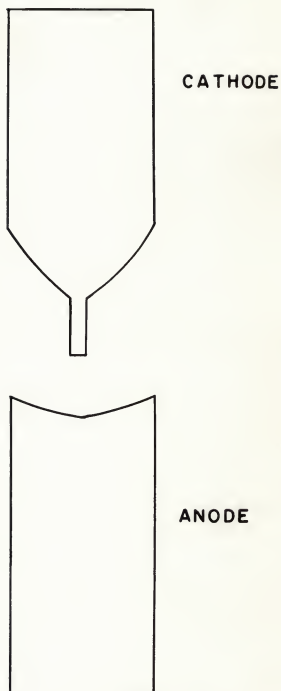


FIG. 2. Cupped and Bullet Shaped Electrodes
used in the D. C. Arc

In the d.c. arc source, the cratered electrodes were used. A 0.1 ml volume of the sample was placed in the crater by means of an Applied Research Laboratories semi-automatic pipette. Each sample was run in triplicate. After the craters were filled, the electrodes were placed in an air oven at 80° C. for one hour and then the temperature was increased to 100° C. for at least five hours. This period of time insured complete drying of the samples on the electrodes. Each electrode was removed from the oven just before being arced, and in this manner there was little possibility of picking up moisture from the air. The solutions which were buffered with lithium tartrate did not form a smooth crust as did the unbuffered solutions. Instead they tended to form a small round ball of residue.

EXPERIMENTAL RESULTS

Effect of Magnesium Upon Intensity of Calcium in the A.C. Spark

The first series of determinations were made to determine the effect of magnesium upon the intensity of calcium in the a.c. spark using unbuffered solutions. The porous cup electrodes were pre-sparked and filled as previously described. The electrodes were then placed in the holders and the electrode spacing adjusted to 2.5 mm with a stainless steel leaf spacer (14). This spacer was mounted on an adjustable shaft so that it could be

moved into place for electrode alignment and then moved back for sparking. The samples were run in duplicate and exposed for sixty seconds. Five plates were exposed giving a total of ten determinations for each sample. After processing the plate, transmittancies were read on the ARL densitometer and the optical density of the calcium line calculated, corrected for calcium in the electrode, and converted to a 0.5 density basis. This correction and conversion was carried out as follows: The optical density of the calcium line in the blank electrode was subtracted from the optical density of the calcium line in the sample. The optical density of the calcium line with no magnesium present was set equal to 0.5, then the optical density of the calcium line when magnesium was present was calculated by means of the ratio, $\frac{\text{O.D. of Ca line} + \text{Mg}}{\text{O.D. pure Ca line}} = \frac{X}{0.5}$ where X is the optical density of the calcium line containing magnesium converted to the 0.5 basis. It is probable that the most erratic results, which in some cases were eliminated, were due to variations in feed rate caused by unequal electrode thickness. These results are shown in Table 2 and Fig. 3.

Effect of Calcium Upon Intensity of Magnesium in the A.C. Spark

The next series of runs were made to determine the effect of calcium upon the intensity of the magnesium line in the a.c. spark using unbuffered solutions. The procedure as given above

Table 2. The effect of magnesium upon the intensity of calcium (3158.9 Å) in the a.c. spark.

Solution	Optical density*						Density	
	Plate number						Standard :	corrected
	22	23	25	26	27	Av.		
Ca 250 ppm	0.650	0.760	0.755	0.751	0.648	0.704	0.0528	0.500
	0.729	0.736	0.614	0.674	0.726			
Ca 250 ppm	0.629	0.630	0.653	0.886	0.607	0.710	0.0935	0.504
Mg 100 ppm	0.729	0.673	0.710	0.845	0.741			
Ca 250 ppm	0.686	0.608	0.678	0.593	0.723	0.684	0.0521	0.486
Mg 200 ppm	0.720	0.696	0.722	0.660	0.757			
Ca 250 ppm	0.753	0.657	0.616	0.643	0.655	0.668	0.0456	0.475
Mg 500 ppm	0.720	0.601	0.663	0.681	0.690			
Ca 250 ppm	0.736	0.632	0.626	0.587	0.658	0.630	0.0635	0.448
Mg 1000 ppm	0.668	0.488	0.649	0.641	0.612			

*All densities have been corrected for electrode impurity.

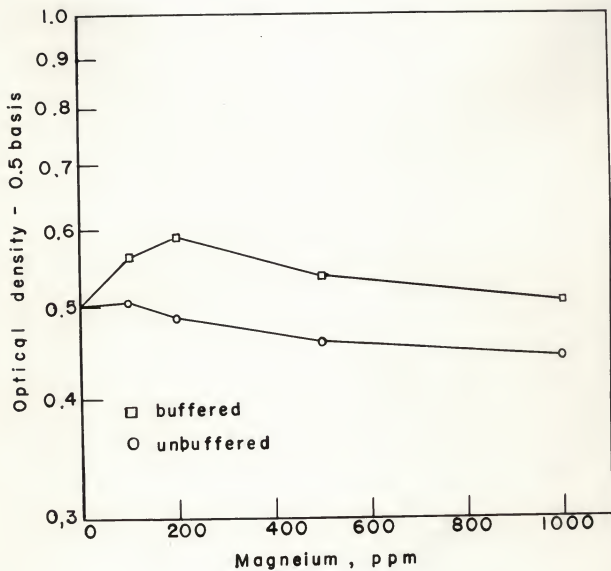


Fig. 3. The Effect of Magnesium upon the Intensity of Calcium (3158.9 Å) in the A.C. Spark.

was followed except that no electrode correction was necessary for the magnesium lines since magnesium was not present as an impurity in the electrodes. The samples were run in triplicate and six plates were exposed giving a total of eighteen analyses on each sample. These results are shown in Table 3 and Fig. 4.

A series of samples were run using the National Carbon Company High Purity Electrodes in order to determine if the electrode impurities had any effect on the calcium and magnesium lines under consideration. The results obtained from these samples indicated that the impurities present in the standard electrodes were not affecting the intensities of the calcium and magnesium spectral lines.

Effect of Calcium Upon Intensity of Magnesium in the D.C. Arc

The effect of calcium upon the intensity of the magnesium line was next determined in the d.c. arc. For this investigation the rotating sector was changed to one-eighth in order that a sixty second exposure could be utilized. This length of exposure was necessary to insure complete consumption of the sample. The results of this series are shown in Table 4 and Fig. 5. Doubtful data were eliminated since occasionally small flakes of the sample were lost in the arc.

Table 3. The effect of calcium upon the intensity of magnesium (2779.8 \AA) in the a.o. spark.

Solution	29	30	31	Optical density				Av.	Standard deviation	: Density : corrected to
				32	33	34	35			
Mg 250 ppm	0.528	0.662	0.656	0.411	0.684	0.609	0.591	0.1035	0.500	
	0.546	0.688	0.714	0.430	0.648	0.611				
	0.428	0.722	0.686	0.456	0.573	0.585				
Mg 250 ppm	0.524	0.417	0.682	---	0.444	0.668	0.614	0.1146	0.519	
	0.656	0.666	0.594	0.428	0.620	0.629				
Ca 100 ppm	0.569	0.832	0.796	0.577	0.660	0.668				
Mg 250 ppm	0.757	0.450	0.639	0.488	0.370	0.516	0.549	0.1242	0.464	
	0.482	0.572	0.830	0.496	0.695	0.558				
Ca 200 ppm	0.596	0.573	---	0.438	0.387	0.492				
Mg 250 ppm	0.620	0.432	0.531	0.376	0.552	---	0.564	0.1079	0.477	
	0.629	0.502	0.678	0.459	0.512	0.606				
Ca 500 ppm	0.745	0.671	0.478	0.660	0.450	0.682				
Mg 250 ppm	0.722	---	0.546	0.372	0.500	0.409	0.539	0.1251	0.457	
	0.460	0.476	0.688	---	0.504	0.593				
Ca 1000 ppm	0.734	0.503	0.733	0.438	0.366	0.577				

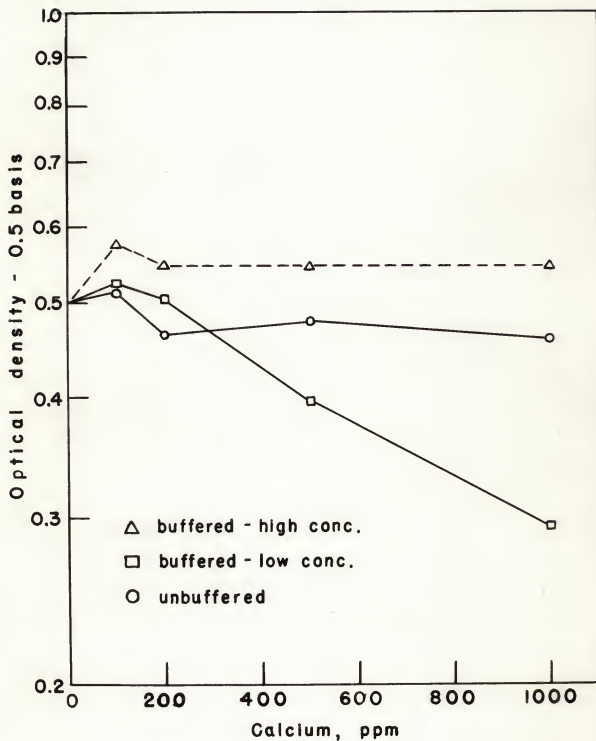


Fig. 4. The Effect of Calcium upon the Intensity of Magnesium (2779.8 \AA) in the A.C. Spark.

Table 4. The effect of calcium upon the intensity of magnesium (2779.8 \AA) in the d.c. arc.

Solution :	Optical density						Av.	Standard deviation :	Density corrected to 0.5 basis		
	41 :			42 :						43 :	44 :
	Plate number										
Mg 250 ppm	0.306 0.454 0.408	0.436 0.446 0.388	0.416 0.404 0.420	0.436 0.412 0.456	0.415	0.0403	0.500				
Mg 250 ppm	0.346 0.404	0.456 0.350	0.462 0.435	0.502 0.465	0.442	0.0614	0.532				
Ca 100 ppm	0.410	0.476	0.546	0.469							
Mg 250 ppm	0.513 0.317	0.432 0.392	0.492 0.584	0.502 0.395	0.450	0.0935	0.542				
Ca 200 ppm	0.294	0.594	0.449	0.434							
Mg 250 ppm	0.445 0.464	0.442 0.564	0.569 0.543	0.318 0.569	0.487	0.0991	0.587				
Ca 500 ppm	0.300	0.569	0.406	0.659							
Mg 250 ppm	0.584 0.660	0.658 0.594	0.780 0.693	0.748 0.568	0.674	0.0749	0.812				
Ca 1000 ppm	0.770	0.600	0.710	0.722							

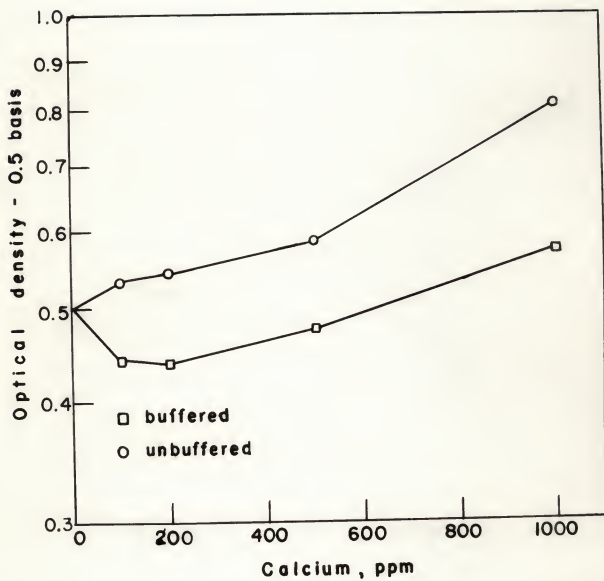


Fig. 5. The Effect of Calcium upon the Intensity of Magnesium (2779.8 Å) in the D.C. Arc

Effect of Magnesium Upon Intensity of Calcium in the D.C. Arc

The d.c. arc source was next utilized to determine the effect of magnesium upon the intensity of the calcium line. The procedure previously described was utilized also in this series. These results are given in Table 5 and Fig. 6.

Selection of Spectroscopic Buffer

In attempting to select a suitable spectroscopic buffer for use with calcium and magnesium, the works of Schrenk and Clements (15) and Langstroth and McRae (12) were considered.

Schrenk and Clements have used lithium tartrate with a good deal of success in this laboratory. The buffer has been used on solutions containing phosphorus, calcium, and potassium. The use of this buffer caused practically complete cancellation of the repression of the phosphorus line by potassium and calcium.

The work of Langstroth and McRae on the transport phenomena of materials in the condensed d.c. spark indicates that a buffer should contain an element of low ionization potential. When such an element is present in sufficient amount it will alter the transport phenomena to such an extent that the extraneous element will not affect it.

An excellent buffer requires that it will form a uniform, tenacious film on the surface of the electrode. This requirement

Table 5. The effect of magnesium upon the intensity of calcium (3158.9 \AA) in the d.c. arc.

Solution	Optical density					Av. : Standard : deviation:0.5 basis	: Density : corrected : to
	Plate number						
	48	49	50	66	67		
Ca 250 ppm	0.372 0.384 ---	0.392 0.473 0.344	0.475 0.512 0.490	0.620 0.577 0.602	0.678 0.699 0.585	0.515 0.1151	0.500
Ca 250 ppm	0.365 ---	0.469 0.342 0.384	0.444 0.406 0.450	0.620 0.648 0.545	0.699 0.620 0.545	0.495 0.1170	0.481
Mg 100 ppm	0.390						
Ca 250 ppm	0.324 0.398 0.404	0.404 0.319 0.340	0.490 0.421 0.426	0.450 0.553 0.745	0.658 0.561 0.668	0.477 0.1321	0.463
Mg 200 ppm							
Ca 250 ppm	0.389 0.429 0.464	0.437 --- 0.437	0.510 0.584 0.389	0.585 0.594 0.569	0.658 0.648 0.553	0.518 0.0929	0.503
Mg 500 ppm							
Ca 250 ppm	0.268 0.234 0.276	0.298 0.278 0.289	0.456 0.446 0.502	0.523 0.585 0.594	0.530 0.620 0.426	0.422 0.1361	0.410
Mg 1000 ppm							

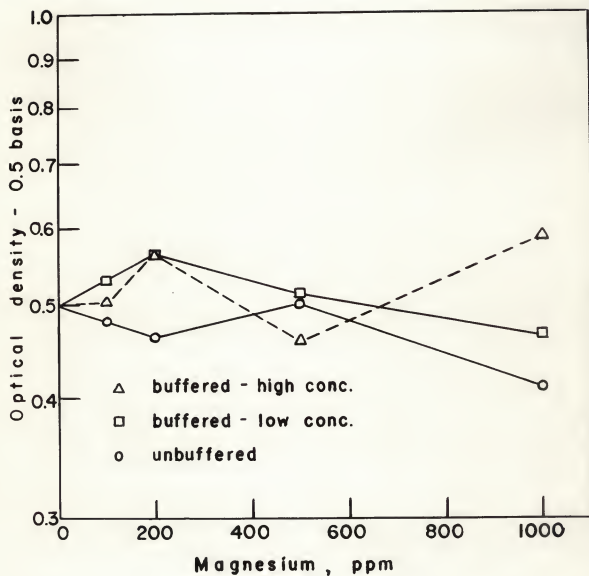


Fig. 6. The Effect of Magnesium upon the Intensity of Calcium (3158.9 Å) in the D.C. Arc

was realized in the case of low concentrations (250 ppm) of sample, but in the higher range (500 to 1000 ppm) the layer was not uniform.

A stock buffer solution containing 81,000 ppm of lithium tartrate was used in compounding the buffered samples for this work. The composition of these samples is shown in Table 1. Two ml of the stock buffer solution were added to each sample of the series for analysis. The concentration of the lithium tartrate in the solutions was approximately 3,860 ppm.

When the buffered samples were used with the porous cup electrodes in the a.c. spark, considerable difficulty was encountered in the passage of the solution through the electrode into the spark. These electrodes had been prepared and treated at the same time as those used with the unbuffered solutions. It was then noted that Ayres and Berg (3) had used pre-arc'd porous cup electrodes and were able to obtain good results. A twenty-five second pre-arc was then applied to the blank electrodes before filling. In this manner the porosity of the electrodes was increased and the samples passed through readily. However, this pre-arcing may influence the determinations since the pre-arcing seemed to enhance the feed rate much more than the pre-sparking.

Table 6 and Fig. 3 show that an enhancement of the calcium line in the a.c. spark occurred when the buffered solutions were used as compared to a repression with the unbuffered solutions. The same phenomenon is true in the d.c. arc as shown in Fig. 6

Table 6. The effect of magnesium upon the intensity of calcium (3158.9 \AA) in the a.c. spark. (Solutions buffered with lithium tartrate).

Solution	Optical density*			Standard deviation	Density corrected to 0.5 basis
	Plate number				
	51	52	53		
Ca 250 ppm	0.509	0.472	0.490	0.552	0.500
	0.840	0.472	0.529		
Ca 250 ppm	0.820	0.671	0.430	0.620	0.562
Mg 100 ppm	0.685	0.594	0.519		
Ca 250 ppm	0.618	0.699	0.582	0.651	0.590
Mg 200 ppm	0.820	0.606	0.582		
Ca 250 ppm	0.399	0.560	0.529	0.560	0.535
Mg 500 ppm	0.631	0.657	0.582		
Ca 250 ppm	0.631	0.455	0.472	0.556	0.504
Mg 1000 ppm	0.657	0.582	0.539	0.0825	

*All densities have been corrected for electrode impurity.

which was plotted from the data tabulated in Table 7.

The effect of the buffer upon the slight repression of the calcium line due to magnesium in the a.c. spark was erratic. The result of a series of three runs in duplicate is shown in Table 8 and Fig. 4.

The enhancement of the magnesium line due to the presence of calcium in the d.c. arc was nullified through the use of the spectroscopic buffer. These results are shown in Table 9 and Fig. 5.

DISCUSSION

Magnesium, when it occurred as an extraneous element, caused a slight depression of the calcium line in the a.c. spark. The average standard deviation over the range of samples studied was approximately 9 percent. The effect of magnesium upon the intensity of calcium is shown in Fig. 3. Buffering the solution with lithium tartrate decreased the slight effect of magnesium upon calcium in the a.c. spark. The standard deviation was 18.7 percent for the buffered solution. When the d.c. arc was employed as an excitation source, the effect of magnesium upon calcium was approximately the same as in the a.c. spark. However, the average standard deviation (24.7 percent) for this series was considerably higher than in the a.c. spark. The spectrographic buffer reacted in the same manner in the d.c. arc as it did in the a.c. spark, except that the standard deviation also was higher (26.7 percent). When the buffer concentration was

Table 7. The effect of magnesium upon the intensity of calcium (3158.9 Å) in the d.c. arc. (Solutions buffered with lithium tartrate).

Solution	Optical density			Standard deviation	Density corrected to 0.5 basis	
	Plate number					
	68	69	70			
Ca 250 ppm	0.444	0.289	0.420	0.0837	0.500	
	0.332	0.332	0.372			
	---	0.426	0.244			
Ca 250 ppm	0.319	0.260	---	0.0934	0.532	
	0.354	0.347	0.377			
	0.516	0.495	0.357			
Mg 100 ppm			0.237			
	Ca 250 ppm	0.367	0.538	0.367	0.0638	0.618
		0.382	---	0.409		
0.502		0.362	0.456			
Mg 200 ppm			0.409			
	Ca 250 ppm	0.403	0.462	0.377	0.0898	0.514
		0.420	0.420	0.292		
0.409		0.272	0.218			
Mg 500 ppm			0.222			
	Ca 250 ppm	0.260	0.292	0.146	0.1345	0.467
		---	0.248	0.314		
0.444		0.420	0.561			
Mg 1000 ppm			0.174			

Table 8. The effect of calcium upon the intensity of magnesium (2779.8 \AA) in the a.g. spark. (Solutions buffered with lithium tartrate).

Solution	Optical density		Av.	Standard deviation	Density corrected to 0.5 basis
	60	61			
Mg 250 ppm	---	0.297	0.313	0.0424	0.500
	0.332	0.292			
Mg 250 ppm	0.409	0.280	0.330	0.0497	0.527
Ca 100 ppm	0.337	0.280			
Mg 250 ppm	0.332	0.252	0.317	0.0538	0.507
Ca 200 ppm	0.398	0.342			
Mg 250 ppm	0.264	0.280	0.249	0.0264	0.398
Ca 500 ppm	0.211	0.229			
Mg 250 ppm	0.208	0.158	0.183	0.0215	0.293
Ca 1000 ppm	0.208	0.184	---		

Table 9. The effect of calcium upon the intensity of magnesium (2779.8 \AA) in the d.c. arc. (Solutions buffered with lithium tartrate).

Solution	Optical density			Standard deviation	Density corrected to 0.5 basis
	63	54	Av.		
Mg 250 ppm	0.903	0.854	0.776	0.0945	0.500
	0.782	---	0.721		
	---	0.538	0.757		
Mg 250 ppm	0.502	0.699	0.824	0.2355	0.442
	---	0.382	1.097		
	0.569	---	0.733		
Mg 250 ppm	0.594	0.553	0.530	0.1170	0.438
	0.810	0.668	0.854		
		0.721	0.710		
Mg 250 ppm	0.824	0.745	0.668	0.1347	0.477
	0.538	---	0.903		
	---	0.629	0.870		
Mg 250 ppm	0.733	0.979	0.699	0.1293	0.574
	0.838	1.000	1.022		
	0.903	0.838	1.022		

increased by a factor of ten, the influence of magnesium on the intensity of the calcium line was not altered appreciably as shown in Fig. 6. Inspection of Fig. 6 also indicates that considerable variation of results was obtained with both concentrations of buffer. The loss of sample during excitation was probably responsible for these variations. Due to the relatively high standard deviations, both in the a.c. spark and the d.c. arc, it is probable that magnesium has very little effect upon the intensity of calcium in either source of excitation.

The intensity of the magnesium line in the a.c. spark decreased slightly when calcium was present as the extraneous element. The decrease in intensity apparently reached a maximum at a concentration of 200 ppm of the extraneous element. The average standard deviation was 20.2 percent for the series. When the spectrographic buffer was added to the solution, the slight decrease in the intensity of the magnesium line was eliminated at low concentrations of the extraneous element. However, when the extraneous element was present at higher concentrations, a marked decrease in the intensity of the magnesium line occurred in the buffered solution. It is probable that this phenomenon may be explained in the following manner: (a) At low concentrations of the extraneous element, the buffer may be present in sufficient quantity to minimize the effect of the extraneous element. (b) At higher concentrations of the extraneous element, the buffer may not be sufficiently concentrated to offset the influence of the extraneous element. For these buffered

solutions the average standard deviation was 13.6 percent. An examination of Fig. 4 would indicate that when an extraneous element is present, a higher concentration of buffer may be required to stabilize the intensity of the magnesium line in the a.c. spark. Therefore, a tenfold increase in buffer was utilized to determine if the influence of calcium upon the intensity of magnesium could be eliminated. The effectiveness of these solutions is shown in Fig. 4 and indicates that it may be possible to obtain a higher precision in determining magnesium under such conditions. The intensity of magnesium was considerably enhanced by calcium when the d.c. arc was used as a source of excitation. When the solution was buffered, the enhancement appeared to be minimized as shown in Fig. 5. The average standard deviations for the unbuffered and buffered solutions were 15.2 percent and 19.3 percent respectively. The data therefore indicate that calcium had very little effect upon the intensity of magnesium in the a.c. spark, but show a considerable enhancement of the same line in the d.c. arc.

Brode and Timma (9) developed an equation which relates the boiling point and excitation potential of an element to the intensity of its spectral line. When the intensities of the calcium and magnesium lines were calculated using the equation proposed by Brode and Timma, the values obtained for the two lines were similar. The difference between these calculated intensities was approximately fifteen percent, which may indicate that the elements should have relatively little effect upon

each other. It is possible that the magnesium and calcium may be converted to their oxides at the temperature obtained in the d.c. arc. If this is true, the energy required to vaporize the oxides would be greater than that of the original material. In the a.c. spark, where the solution is fed directly into the spark gap, it is improbable that the magnesium and calcium would be converted to their oxides. Therefore, the effect of one element upon another should be different in the two sources of excitation, and the results of this research has shown this to be true.

CONCLUSIONS

It has been shown that calcium and magnesium can influence each other in both the a.c. spark and the d.c. arc. These effects are summarized as follows: (a) The presence of magnesium will cause a decrease in the intensity of the calcium line in the a.c. spark and also in the d.c. arc. The results in the a.c. spark tend to be more uniform than those obtained in the d.c. arc; (b) When calcium is present as an extraneous element in the determination of magnesium, there will be a slight repression of the magnesium line in the a.c. spark and a considerable enhancement in the d.c. arc.

The effects stated above have been eliminated somewhat through the use of lithium tartrate as a spectrographic buffer. This buffer minimized the effect of magnesium upon calcium. The results were erratic when a low concentration of buffer was used

to eliminate the effect of calcium upon magnesium. However, higher concentrations of buffer have minimized the effect of calcium upon the intensity of the magnesium line.

The standard deviations as indicated in Tables 2 through 9 are high. This would tend to indicate several sources of error, which in all probability, could be partially eliminated through a more comprehensive study of the effects of these elements upon each other.

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THE INFLUENCE OF EXTRANEOUS ELEMENTS ON
LINE INTENSITIES IN THE A.C. SPARK AND THE D.C. ARC

by

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This research was conducted to study the influence of extraneous elements on line intensities in the a.c. spark and the d.c. arc. Calcium and magnesium were the investigated elements. A Bausch and Lomb, large, Littrow spectrograph with quartz optics was used throughout the research.

Kodak Number 33 photographic plates were found to give good results when calcium and magnesium were investigated. All plates were developed four minutes in dilute Kodak X-ray developer and fixed in X-ray fixer. The transmittances of the lines were measured with a projection type Allied Research Laboratories - Dietert densitometer.

The concentration of the element being investigated was held constant at 250 ppm. The extraneous element was varied as follows: 0; 100; 200; 500; and 1000 ppm. Lithium tartrate was used as a spectrographic buffer to minimize the influence of the extraneous elements. When a high concentration of buffer was employed, the concentration of the element under investigation was 125 ppm.

Results indicated that the presence of magnesium will cause a decrease in the intensity of the calcium line in the a.c. spark and also in the d.c. arc. When calcium is present as an extraneous element in the determination of magnesium, there will be a slight repression of the magnesium line in the a.c. spark and a considerable enhancement in the d.c. arc. These effects have been eliminated somewhat through the use of lithium tartrate as a spectrographic buffer. This buffer minimized the effect of

magnesium upon calcium. Erratic results were obtained when a low concentration of buffer was used to eliminate the effect of calcium upon magnesium. However, higher concentrations of buffer have minimized the effect of calcium upon the intensity of the magnesium line in the a.c. spark.

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